

Day of Week	Date	Skill	Plan
M	9/29/2014	Written objectives located at bottom of lesson plan Psc. 2.1.4, 2.2.1-2 acceleration	Entry: Review based on test grades: Prof Penguin: https://www.youtube.com/watch?v=MYuh5yErdFA Review quiz: Test Students to look for their mistakes Discuss: Classwork/Homework: "An Eye on the Ion" & Hogwartz Periodic Table Practice atomic structure review from Passing EOC Ch12, p193-205 DO NOT write in books.
T	9/30/2014	Club Day Psc. 2.1.4, 2.2.1-2	Entry: Review Homework: Review work Quiz/Test: Discuss: Ch 19 lesson 3.2-3.3 ions symbols and charges, oxidation numbers Classwork/Homework: Wkst Valence electrons, Metals, nonMetals and Oxidations assess while working with students
W	10/1/2014	Psc. 2.1.4, 2.2.1-2	Entry: Begin Study guide on Periodic Table: bonds [front side] Review Homework: Quiz/Test: Quiz on Lewis Dot structures, Ion notation and Ionic Bonding SONG: ionic & Covalent Bonds https://www.youtube.com/watch?v=QlFTT-_xLo Discuss: Naming Compounds, polyatomic ions & diatomic molecules Classwork/Homework: Types of Chemical Bonds & Oxidation Practice assess while working with students
H	10/2/2013	Psc. 2.1.4, 2.2.1-2	Entry: Continue study guide [back side] Review Homework: Naming Compounds Types of Chemical bonds Quiz/Test: Discuss: Covalent bonding Classwork/Homework: Practice Covalent Bonds
F	10/3/2013	PSc. 2.2.2-5	Entry: periodic table worksheet Review Homework: Covalent bonds practice Quiz/Test: Discuss: Naming Chemical Compounds Classwork/Homework: Practice Naming Chemical Compounds Extra: Discuss: Writing Chemical Formulas; Practice Writing Chemical Formulas Single & double replacements- http://www.youtube.com/watch?v=4T9IAOeb0qo PSc.2.1.4 Interpret the data presented in the Bohr model diagrams and dot diagrams for atoms and ions of elements 1 through 18.

Standards PSc:	<p>2.1.4 Interpret the data presented in the Bohr model diagrams and dot diagrams for atoms and ions of elements 1-18.</p> <p>2.2.1 Infer valence electrons, oxidation numbers, and reactivity of an element based on its location in the Periodic Table.</p> <p>2.2.2 Infer type of chemical bond that occurs, whether covalent, ionic, metallic, in a given substance.</p> <p>2.2.3 Predict chemical formulas and names for simple compounds based on knowledge of bond formation and naming conventions.</p> <p>2.2.4 Exemplify the law of conservation of mass by balancing chemical equations.</p> <p>2.2.5 Classify types of reactions such as synthesis, decomposition, single replacement, or double replacement.</p>
Essential Q's #:	How is the composition of atoms related to chemistry in general?
Unit Questions #:	<p>3. What is involved in a chemical reaction and why do they need to be balanced?</p> <p>4. Why do atoms want/need to bond?</p> <p>5. How do I name a chemical and what can I learn from the name?</p> <p>6. How do I write a chemical formula and what can I learn from writing the formula?</p> <p>7. How can I use the periodic table to tell me everything I need to know about each element (atomic #, proton #, neutron #, atomic mass, valence #.....)?</p>
Content Q's #:	<p>2. What are the components of an atom?</p> <p>3. What are oxidation numbers?</p> <p>4. How do I use the periodic table?</p>
Learning Objectives	<p>2. I will understand and be able to draw Bohr models and dot diagrams.</p> <p>3. I will understand periodic law and how to utilize the periodic table for oxidation numbers, valence electrons, number of energy levels, classification between metals and nonmetals, and the type of bond.</p> <p>4. I will be able to distinguish the difference between atoms and their isotopes.</p>

